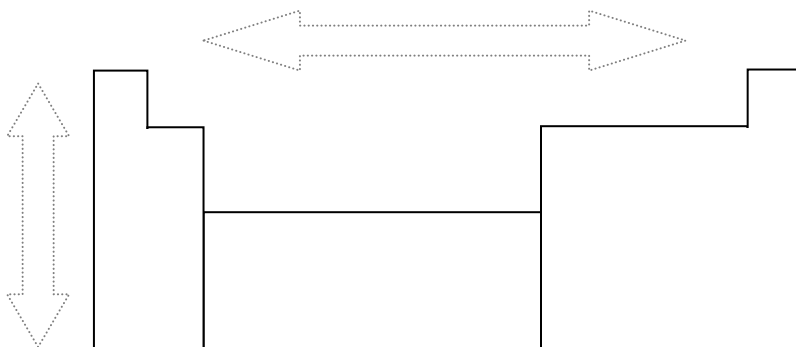




Periodicity

Complete the following information on different types of periodic trends using your textbook. First, define the concept of the trend. Then, explain the trend for both groups and periods on the periodic table (be sure to explain why!). Also shade in the arrows on the periodic tables to **label whether the trend is increase or decreasing in that direction.**

Atomic Number



Define

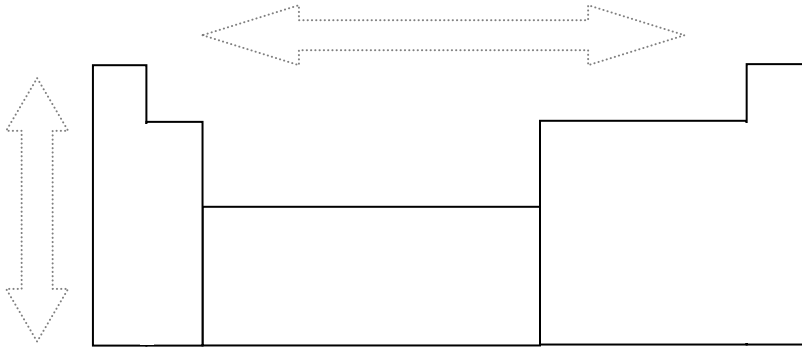
Periodic Trend

Why?

Group Trend

Why?

Atomic Radius



Define

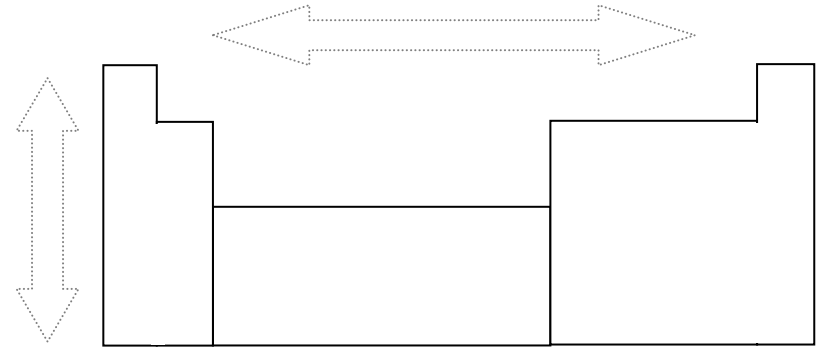
Periodic Trend

Why?

Group Trend

Why?

Ionic Radius



Define

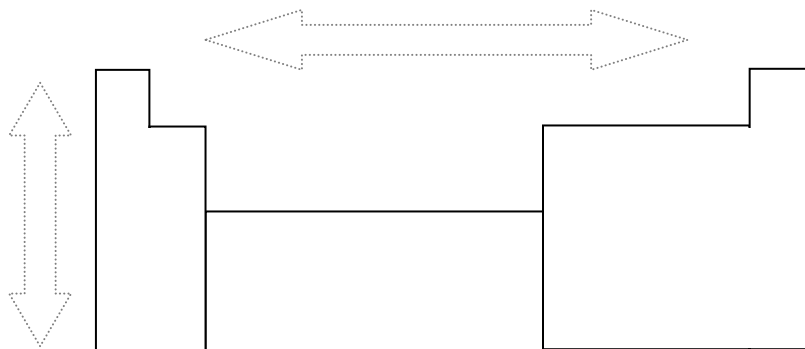
Periodic Trend

Why?

Group Trend

Why?

Ionization Energy



Define

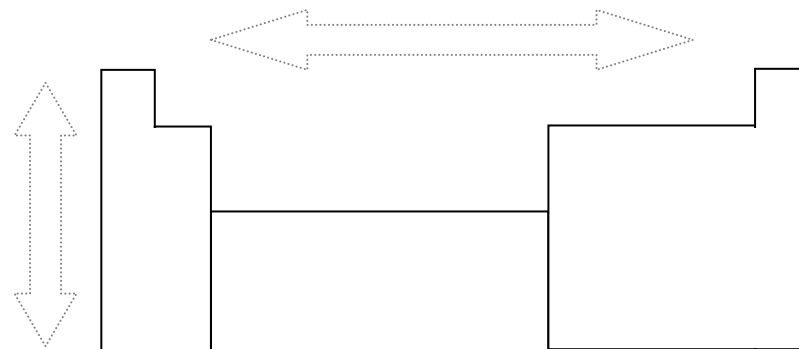
Periodic Trend

Why?

Group Trend

Why?

Electronegativity



Define

Periodic Trend

Why?

Group Trend

Why?

Concept Questions

1. Explain the octet rule.
2. How is ionization energy related to the rules used to predict the formation of ions (the octet rule)?
3. Why is it more difficult to remove the second or third electron?
4. Complete the Section 3 review (p. 169 #20-23).